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**Teacher Guide -  
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In this simulation, students will investigate three of the fundamental gas laws, including Boyle's Law, Charles' Law and Gay-Lussac's Law. Students will have the opportunity to visually examine the effect of changing the associated variables of pressure, volume, or temperature in each situation.

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## **POGIL Chemistry Teachers Edition**

Using Active Folders in Your Classroom Meeting Objectives National and state science standards provide the focus for each folder. Specific objectives for each folder are listed on the teacher page.

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## **Holt McDougal Modern Chemistry Chapter 10: States of ...**

Total Support for  
*Page 7/25*

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Teachers. Student Edition and a comprehensive Teacher's Edition are available in print and digital formats. Active Chemistry Learning Community provides teachers with resources to prepare lessons as well as share and compare with other teachers in an online community.

**AP\* Chemistry**  
**GASES**

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## Chemistry Gases

### Teacher Edition

Lesson Plans

Chemistry: Matter and  
Change • Chapter 14

59 The Ideal Gas Law

LESSON PLAN pages

434–439 1 class

session(s) 14.3 Key: SE

Student Edition, TWE

Teacher Wraparound

Edition, Section

Objectives • Relate the

amount of gas present

to its pressure,

temperature, and

volume by using the

ideal gas law. •

Compare the

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properties of real and  
ideal gases.

## **Classroom Resources | Gases | AACT**

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**chapter 5 ap**  
*Page 10/25*

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**chemistry zumdahl  
Flashcards and ... -  
Quizlet**

Put on the goggles.  
Have students make  
the hydrogen gas:  
Place a small piece  
(1-2 cm) of magnesium  
ribbon into a test tube.  
Add several drops of  
hydrochloric acid to the  
test tube. Once it starts  
to bubble, the test can  
be conducted. Light a  
splint and place it near  
the opening of the test  
tube.

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Flashcards and  
Study Sets | Quizlet**

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- **BOYLE'S LAW:** "father of chemistry" --the volume of a confined gas is inversely proportional to the pressure exerted on the gas. ALL GASES BEHAVE IN THIS MANNER!
- Robert Boyle was an Irish chemist. He studied PV relationships using a J-tube set up in the multi-story entryway of his home.

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**Teacher's edition:  
Various ...**

temperature pressure  
volume In your  
textbook, read about  
the combined gas law  
and Avogadro's  
principle. Circle the  
letter of the choice that  
best completes the  
statement or answers  
the question. 7. The  
variable that stays  
constant when using  
the combined gas law  
is d. volume. b.  
pressure. a. amount of

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gas. c. temperature.

**Chemistry Study  
Guide for Gases**

Douglas Fisher, PhD, is a Professor in the Department of Teacher Education at San Diego State University. He is the recipient of an International Reading Association Celebrate Literacy Award as well as a Christa McAuliffe award for Excellence in Teacher Education. He

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has published

**Chemistry Science  
Notebook: Student  
Edition**

A device to measure atomic pressure. A device for measuring the pressure of a gas in a container. A unit of pressure measurement.

Barometer A device to measure atomic pressure. Manometer A device for measuring the pressure of a gas in



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a container. Form of matter wherein substance that uniformly fills any cont....

## **Classroom Resources | Gas Laws Simulation | AACT**

The ideal gas law, also known as the combined gas law, is a combination of all the variables in the previous gas laws. The ideal gas law is

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expressed by the formula  $PV = nRT$  where  $P$  = pressure  $V$  = volume  $n$  = number of moles of gas  $R$  = ideal gas constant  $T$  = absolute temperature The value of  $R$  depends on the units of pressure, volume and temperature.

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Study Guide for  
*Page 18/25*

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Content Mastery  
Answer Key Chemistry:  
Matter and Change  
T195 Name Date Class  
76 Chemistry: Matter  
and Change • Chapter  
13 Study Guide for  
Content Mastery  
Section 13.3 Liquids  
and Solids In your  
textbook, read about  
liquids and solids. In  
the space at the left,  
write true if the  
statement is true; if the  
statement is false,

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**The Gas Laws pages  
419 - Glencoe**

Chemistry Lesson #11

The Property of Gases

o We know that all gas particles are not the same size - a hydrogen molecule only has two protons and two electrons - where as a molecule of chlorine gas has 34 protons, 36 neutrons plus 34 electrons. Thus when Avogadro purposed that equal volumes of gases at the same

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temperature

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**Chemistry Lesson  
Plans #11 - The  
Property of Gases**

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Microscale Gas

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Chemistry. For classroom use by chemistry teachers. This website provides instructions for the generation of gases on a microscale level inside 60 mL syringes along with instructions for chemical demonstrations and student laboratory experiments with the gases.

## **Chemistry Chapter 14 Retest**

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**Assignment**

First the teacher will burn a small sample of propane gas in a beaker. Next the teacher will burn a small sample of methane gas. Students will create particle diagrams in order support their explanation and model their observations as they improve their understanding of gas density.

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**Study Guide for  
Content Mastery -  
Teacher Edition -  
Chemistry**

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The course is wrapped



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up with a discussion of the ideal gas laws.

Please note that this is Volume 2 of the Teacher's Edition.

Volume 1 of the Teacher's Edition, the Student Edition and the Manipulative Set must be purchased separately to have all necessary materials to complete this course.