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5 The Periodic Law - Jefferson Township Public Schools

Holt Modern Chemistry Review CHAPTER 5: THE PERIODIC LAW The following pages contain the bulk (but not all) of the information for the chapter 5 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.

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7 Chemical Formulas and Chemical Compounds

These lecture presentations were designed for my high school Chemistry I Honors class. Students of high school and college general chemistry may find them useful as a supplement to their own class notes or as a review. Teachers, please feel free to use and modify them for your own classes.

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Write empirical formulas to match the following molecular formulas: CH₃O₂ a. C₂H₆O₄N₂O₅ b. N₂O₅HgCl c. Hg₂Cl₂CH₂ d. C₆H₁₂ 2. C₄H₈A certain hydrocarbon has an empirical formula of CH₂ and ...

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CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

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CHAPTER 6 REVIEW Chemical Bonding SECTION 5 SHORT ANSWER Answer the following questions in the space provided. 1. Identify the major assumption of the VSEPR theory, which is used to predict the shape of atoms. Pairs of valence electrons repel one another. 2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn't water a linear ...

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6 Chemical Bonding

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